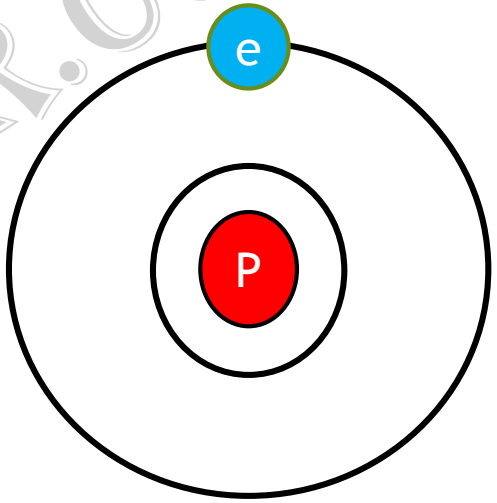
PERIODIC PROPERTIES OF ELEMENTS FROM HYDROGEN TO NEON

By Tutoring Masters

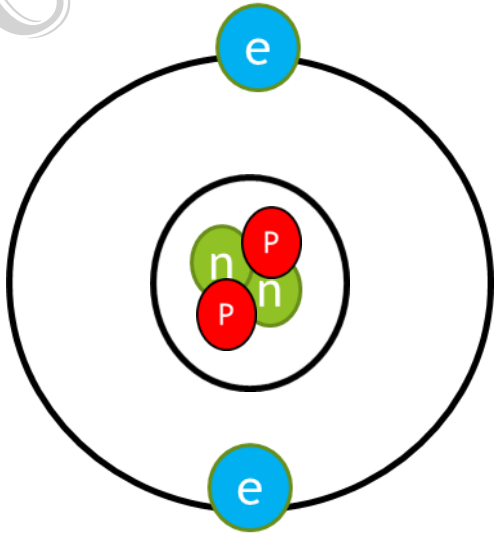
HYDROGEN

- Symbol H
- Atomic number: 1
- Mass number: 1
- Atomic mass: 1.00784u
- Nature & state: non-metal, diatomic gas
- Electronic configuration: $1S^1$
- Melting point: -259.14^oC
- Boiling point: -252.87^oC
- Group number: I-A
- Ions $H_30^+(H^+)$, H^-



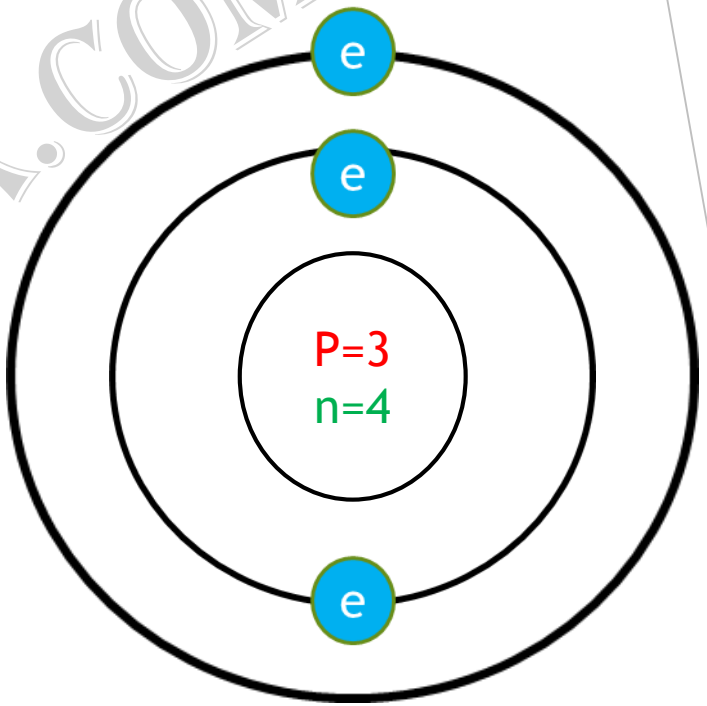
HELIUM

- Symbol He
- Atomic number: 2
- Mass number: 4
- Atomic mass: 4.002602u
- Nature & state: Noble gas
- Electronic configuration: $1S^2$
- Melting point: $-272.2^{O}C$
- Boiling point: $-268.9^{O}C$
- Group number: VIII-A
- Oxidation state: 0



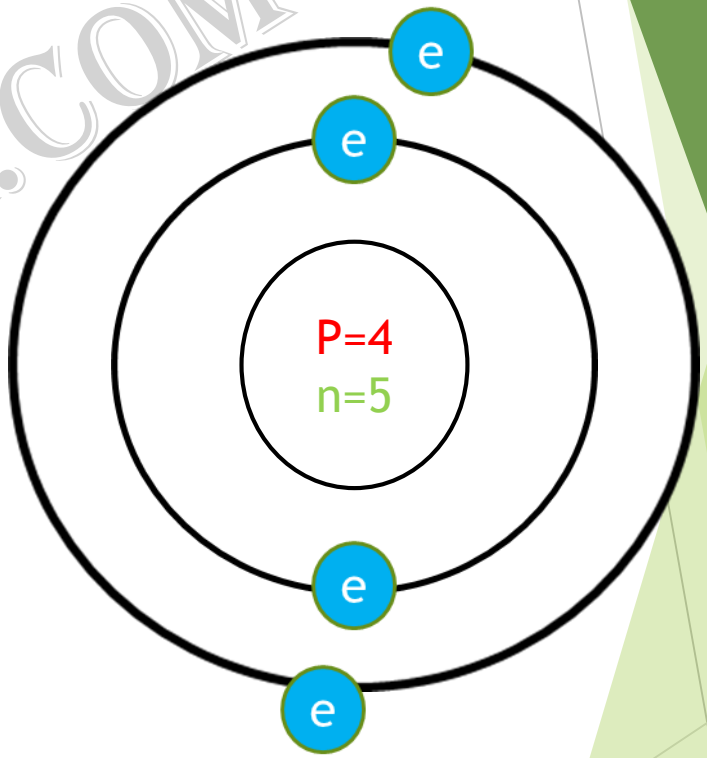
LITHIUM

- Symbol Li
- Atomic number: 3
- Mass number: 7
- Atomic mass: 6.941u
- Nature & state: Metal, solid
- Electronic configuration: $1S^2$, $2S^1$
- Melting point: 180.54^oC
- Boiling point: 1342^oC
- Group number: I-A (Alkali metal)
- Ion: Li^+



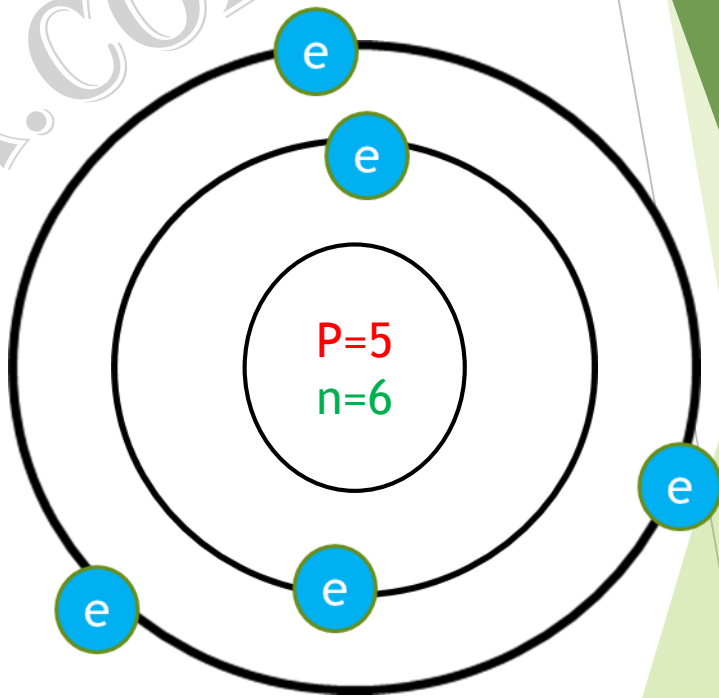
BERYLLIUM

- Symbol Be
- Atomic number: 4
- Mass number: 9
- Atomic mass: 9.012u
- Nature & state: Metal, solid
- Electronic configuration: $1S^2, 2S^2$
- Melting point: 1287^oC
- Boiling point: 2970^oC
- Group number: II-A (Alkaline-earth metal)
- Ion: Be^{+2}



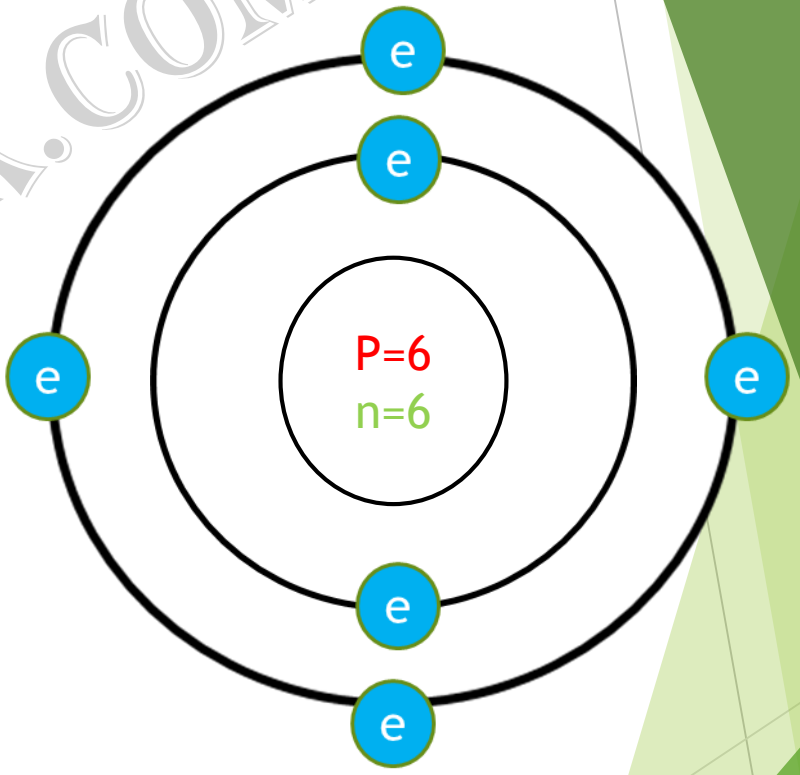
BORON

- Symbol B
- Atomic number: 5
- Mass number: 11
- Atomic mass: 10.811u
- Nature & state: Metalloid, solid
- Electronic configuration: $1S^2$, $2S^2$, $2P^1$
- Melting point: 2300^oC
- Boiling point: 2550^oC
- Group number: III-A (Boron family)
- Ion: B^{+3}



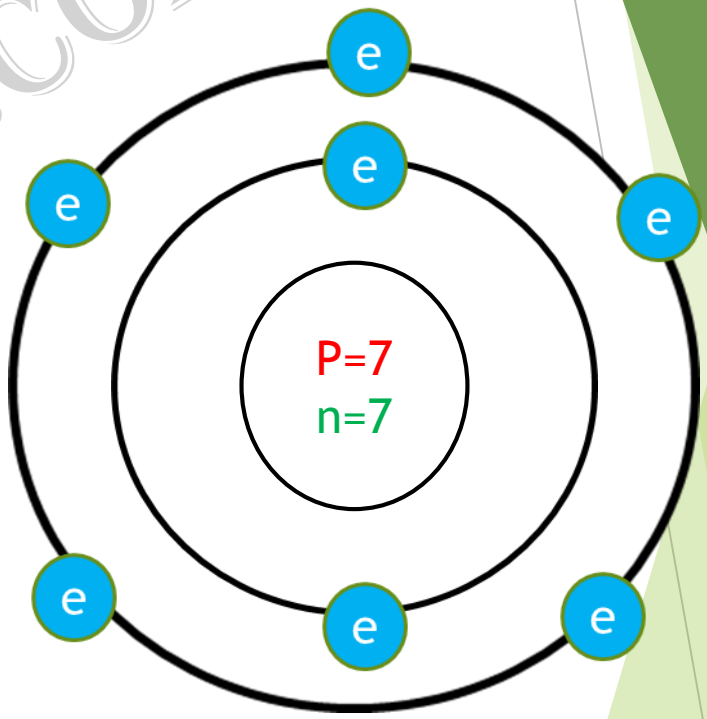
CARBON

- Symbol C
- Atomic number: 6
- Mass number: 12
- Atomic mass: 12.0107u
- Nature & state: Non-metal, solid
- Electronic configuration: $1S^2$, $2S^2$, $2P^2$
- Melting point: $3,367^oC$
- Boiling point: 4827^oC
- Group number: IV-A (Carbon family)
- Ion: C^{+4} ,C^{-4} ,C^{+2}



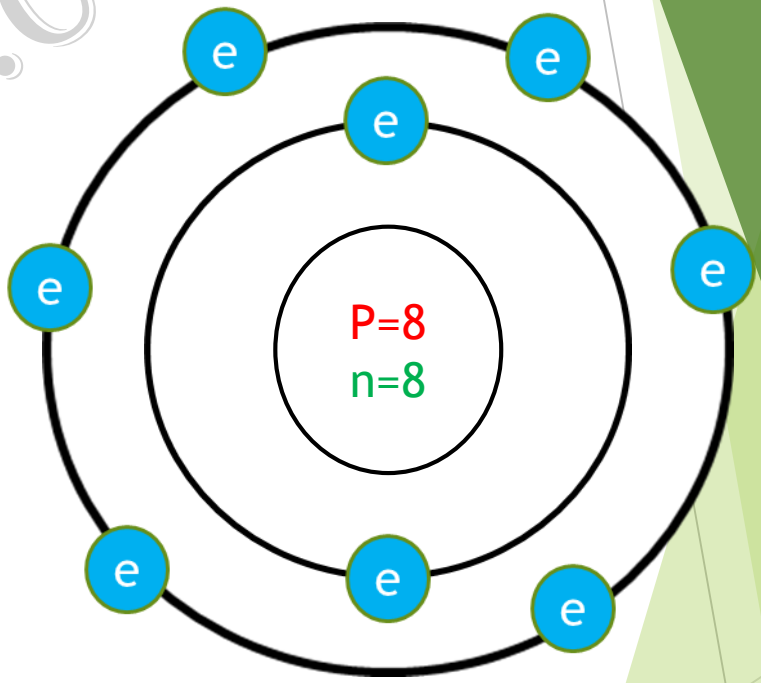
NITROGEN

- Symbol N
- Atomic number: 7
- Mass number: 14
- Atomic mass: 14.0067u
- Nature & state: Non-metal, diatomic gas
- Electronic configuration: $1S^2$, $2S^2$, $2P^3$
- Melting point: -210^oC
- Boiling point: -195.8^oC
- Group number: V-A (Nitrogen family)
- Ion: N^{+3} ,N^{+4}, N^{+5},N^{+1},N^{-1},N^{-3}



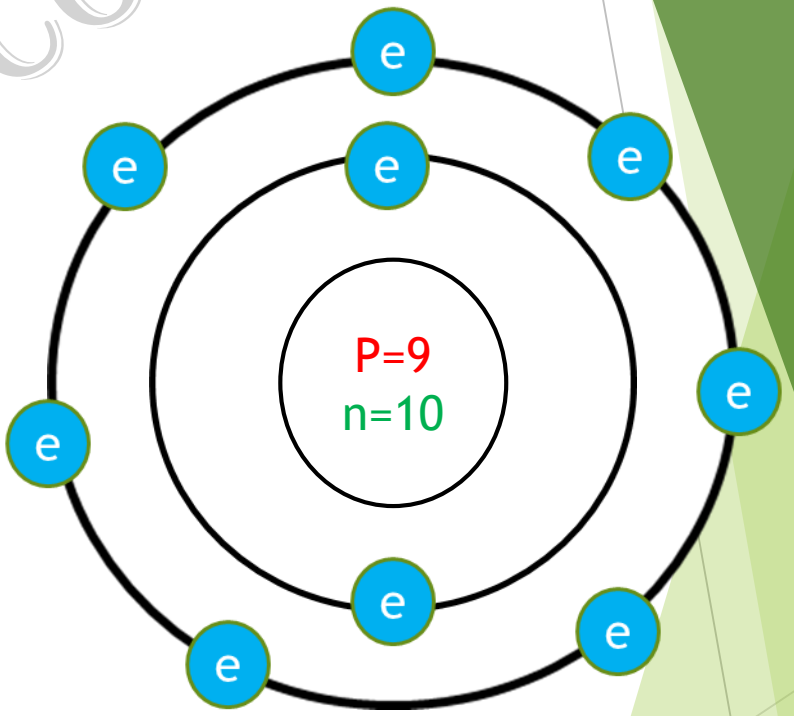
OXYGEN

- Symbol O
- Atomic number: 8
- Mass number: 16
- Atomic mass: 15.999u
- Nature & state: Non-metal, diatomic gas
- Electronic configuration: $1S^2, 2S^2, 2P^4$
- Melting point: -218.8^oC
- Boiling point: -183^oC
- Group number: VI-A (Oxygen family)
- Ion: O^{-2} ,O^{-1}, O^{+2}



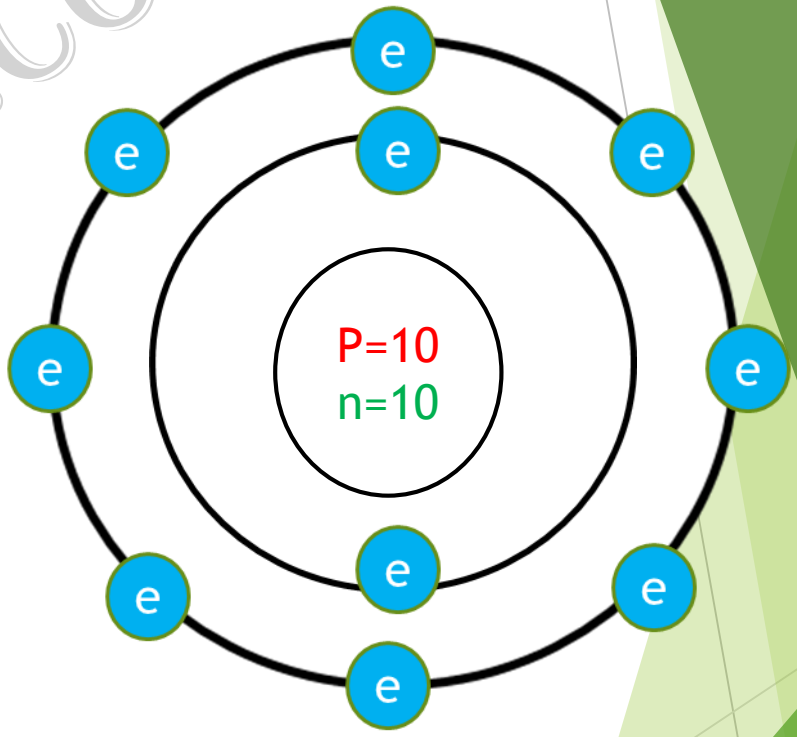
FLUORINE

- Symbol F
- Atomic number: 9
- Mass number: 19
- Atomic mass: 18.9984u
- Nature & state: Non-metal, diatomic gas
- Electronic configuration: $1S^2, 2S^2, 2P^5$
- Melting point: $-219.6^{O}C$
- Boiling point: $-188.1^{O}C$
- Group number: VII-A (Halogen)
- Ion: F^{-1}



NEON

- Symbol Ne
- Atomic number: 10
- Mass number: 20
- Atomic mass: 20.1797u
- Nature & state: Non-metal, noble gas
- Electronic configuration: $1S^2, 2S^2, 2P^6$
- Melting point: $-248.6^{O}C$
- Boiling point: $-246^{O}C$
- Group number: VIII-A
- Oxidation state: 0



Thanks for watching

Coming up: Electronic configuration of the next 108 elements!!!!!